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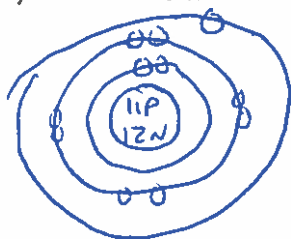
SNC 1PI
Chemistry Practice Exam

1. Complete the table below for each neutral atom.

Element Name	Atomic #	Mass #	# of protons	# of neutrons	# of electrons
beryllium	4	9	4	5	4
argon	18	40	18	22	18
fluorine	9	19	9	10	9
phosphorus	15	31	15	16	15
sodium	11	23	11	12	11

2. Draw Bohr-Rutherford diagrams for the following:

(a) sodium atom



(b) nitrogen



3. Complete the missing sections in the following table.

Name	Formula
magnesium oxide	MgO
sodium fluoride	NaF
aluminum nitride	AlN
potassium sulfide	K ₂ S
lithium iodide	LiI
lithium oxide	Li ₂ O
aluminum chloride	AlCl ₃
magnesium sulfide	MgS
calcium oxide	CaO
potassium bromide	KBr

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4. Complete the following chart:

Name of Chemical Family	Alkali Metals	Halogens	Noble Gases
Location on the Periodic Table (Group Number)	1	17 (7)	18 (8)
Number of Valence Electrons	1	7	8
Physical Properties	silver soft not dense	all three states non metals	odorless colourless
Chemical Properties	highly reactive	highly reactive	not reactive

5. Give ten examples of physical properties.

colour

texture

colour

odour

melting point

boiling point

density

solubility

polarity

density

6. What are the 5 signs of a chemical change?

bubbles (gas)

colour change

heat / light given off

difficult to reverse

precipitate