

Molecular Compounds

- _____ are compounds consisting of _____
- they are formed when atoms _____ electrons to form stable arrangements
- this sharing of electrons forms a _____ (two nonmetal atoms sharing electrons)
- molecular compounds are non-electrolytes (they do not conduct electricity when dissolved in water)

Type of compound	Molecular	Ionic
Atoms involved		
Conduction		
Bond type		

Writing Formulas For Molecular Compounds

- The prefix in the name becomes the subscript in the formula
 - a. sulfur dibromide
 - b. carbon tetrafluoride
 - c. phosphorus pentaxide
 - d. dinitrogen trioxide

Naming Molecular Compounds

- check to make sure that both elements are _____
- write the names of both elements in the _____ as the formula
- replace the ending of the _____ element with _____
- add _____ – mono is never used for the first element

Prefix	#
mon(o)	
di	
tri	
tetra	
penta	

- a. CO
- b. CS₂
- c. PBr₅
- d. SO₃
- e. CF₄

Common Names of Some Molecular Compounds

Water

Ammonia

Nitric Oxide

Methane

Ozone

Diatomic Elements

- Elements which exist in nature in pairs (never on their own)
- Super 7
 - seven of them
 - they form the number seven

Element	Symbol	Molecule
hydrogen	H	
nitrogen	N	
oxygen	O	
fluorine	F	
chlorine	Cl	
bromine	Br	
iodine	I	