

SNC 1D1
Chemistry Practice Exam

1. Complete the table below for each neutral atom.

Element Name	Atomic #	Mass #	# of protons	# of neutrons	# of electrons
	4				5
		40		18	
fluorine					
	15				
			11		12

2. Draw Bohr-Rutherford diagrams for the following:

(a) sodium atom

(b) nitrogen

(c) stable ion formed by oxygen

(d) stable ion formed by calcium

3. Complete the missing sections in the following table.

Name	Formula
potassium sulfide	
magnesium nitride	
	Mg_3P_2
	$FeBr_3$
copper (II) sulfide	
aluminum nitride	
	CaO
	NiO
calcium bromide	
	BeF

4. Draw structural diagrams for each of the following compounds.

a) potassium sulfide

b) methane (CH_4)

c) carbon monoxide