

SNC 1D1
Chemistry Practice Exam

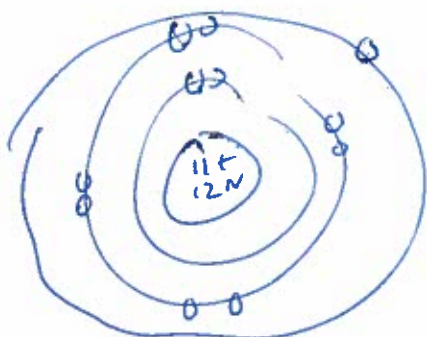
1. Complete the table below for each neutral atom.

Element Name	Atomic #	Mass #	# of protons	# of neutrons	# of electrons
beryllium	4	9	4	5	4
argon	18	40	18	18	22
fluorine	9	19	9	9	10
phosphorus	15	31	15	15	16
sodium	11	23	11	11	12

electrons neutrons

2. Draw Bohr-Rutherford diagrams for the following:

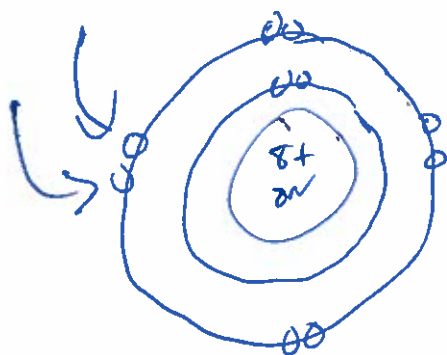
(a) sodium atom



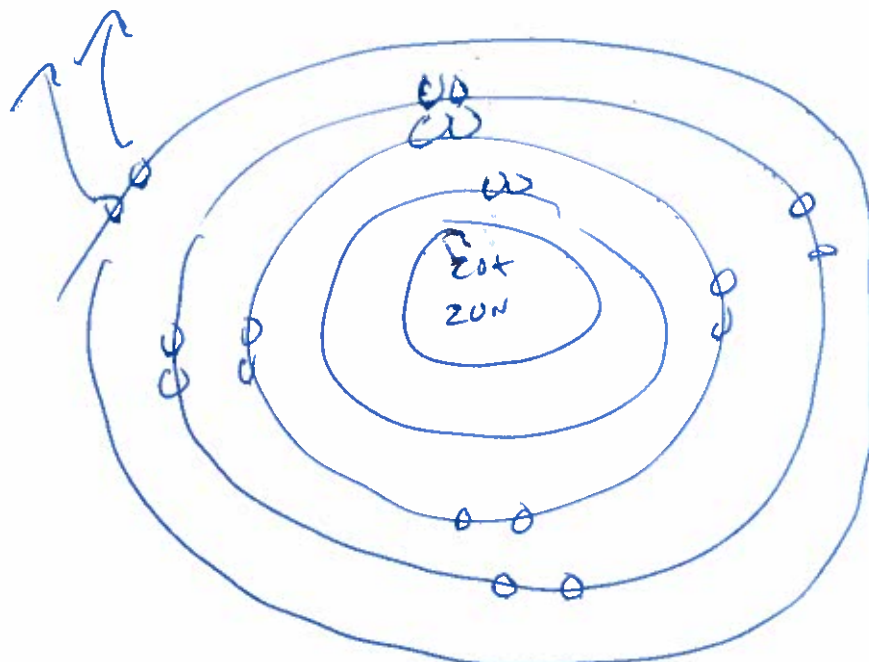
(b) nitrogen



(c) stable ion formed by oxygen



(d) stable ion formed by calcium



3. Complete the missing sections in the following table.

Name	Formula
potassium sulfide	K_2S
magnesium nitride	Mg_3N_2
magnesium phosphide	Mg_3P_2
iron (III) bromide	$FeBr_3$
copper (II) sulfide	CuS
aluminum nitride	AlN
calcium oxide	CaO
nickel (II) oxide	NiO
calcium bromide	$CaBr_2$
beryllium fluoride	BeF_2

4. Draw structural diagrams for each of the following compounds.

a) potassium sulfide



b) methane (CH_4)



c) carbon monoxide

