

**Molecular Compounds**

- molecular compounds are compounds consisting of \_\_\_\_\_
- they are formed when atoms \_\_\_\_\_ electrons to form stable arrangements
- this sharing of electrons forms a \_\_\_\_\_ (two nonmetal atoms sharing electrons)
- molecular compounds are \_\_\_\_\_ (they do not conduct electricity when dissolved in water)

Type of compound	Molecular	Ionic
Atoms involved		
Conduction		
Bond type		

**Writing Formulas For Molecular Compounds**

- The prefix in the name becomes the subscript in the formula
  - a. sulfur dibromide \_\_\_\_\_
  - b. carbon tetrafluoride \_\_\_\_\_
  - c. phosphorus pentaxide \_\_\_\_\_
  - d. dinitrogen trioxide \_\_\_\_\_

**Naming Molecular Compounds**

- check to make sure that both elements are \_\_\_\_\_
- write the names of both elements in the \_\_\_\_\_ as the formula
- replace the ending of the \_\_\_\_\_ element with \_\_\_\_\_
- add \_\_\_\_\_ – mono is never used for the first element

Prefix	#
mon(o)	
di	
tri	
tetra	
penta	

- a. CO \_\_\_\_\_
- b. CS<sub>2</sub> \_\_\_\_\_
- c. PBr<sub>5</sub> \_\_\_\_\_
- d. SO<sub>3</sub> \_\_\_\_\_
- e. CF<sub>4</sub> \_\_\_\_\_

